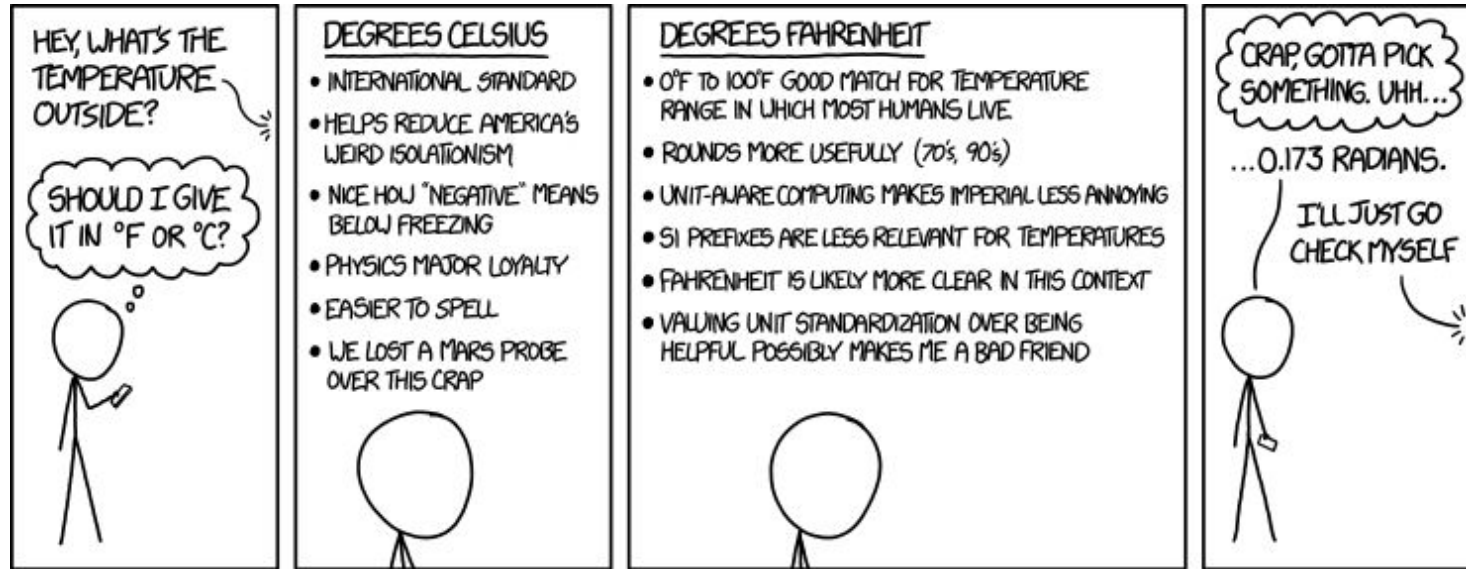


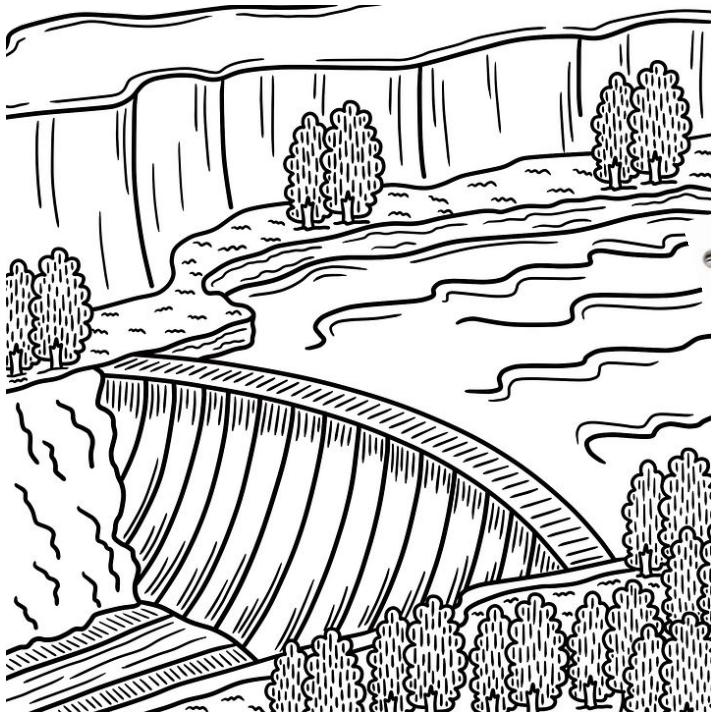
# Absolute Temperature???

- Angad Singh, 3<sup>rd</sup> year BSMS



Powered by Fundamentals of Statistical and Thermal Physics, by F.Reif  
~ For a summer reading project under Prof.Deepak Dhar

# Thermodynamics: *Good old Thermometers*



Any small macroscopic system can be used to compare other larger 'Reservoirs' by comparing a linked macroscopic parameter.

Not to scale\*

# Most Probable Distribution

$\Omega$  is the count of the number of 'states' a system could be in, given what we know about its macroscopic parameters, like Length, Pressure, Magnetization, Energy etc.

$\Omega(E) = \#$  of possible states of a system having an energy between  $E$  and  $E+\delta E$ .

Now consider two systems (A and B) in contact with each other and nothing else

$$(E_A + E_B = E_{\text{tot}})$$

We can get the states available for a particular internal distribution simply by multiplying their individual  $\Omega(E)$  counts.

$$\Omega_0(E, E_{\text{tot}} - E) = \Omega_A(E) \Omega_B(E_{\text{tot}} - E)$$

# Entropy: Probability in Disguise

$$\beta \stackrel{\text{def}}{=} \partial \text{Ln}(\Omega) / \partial E$$

For a ‘typical’ system

$$1/\beta \stackrel{\text{def}}{=} kT$$

$$\Omega(E) \propto (E - E_0)^f$$

$$S \stackrel{\text{def}}{=} k \text{Ln}(\Omega)$$

$$kT = 1/(\partial S / \partial E) \propto (E - E_0)/f$$

Some sense of average energy per ‘particle’

$$1/T = \partial S / \partial E$$

# *Non-Typical Systems*

What if the number of states don't increase very very exponentially with more energy?

Introducing... Spin Systems! [You do have to forget that Kinetic energy exists]

Imagine a large number (N) of weakly interacting magnetic moments in a system, and then put an identical copy next to it.

Each system now has an upper cap on energy, and only  $2^N$  states overall.

$$N! / n!(N-n)!$$

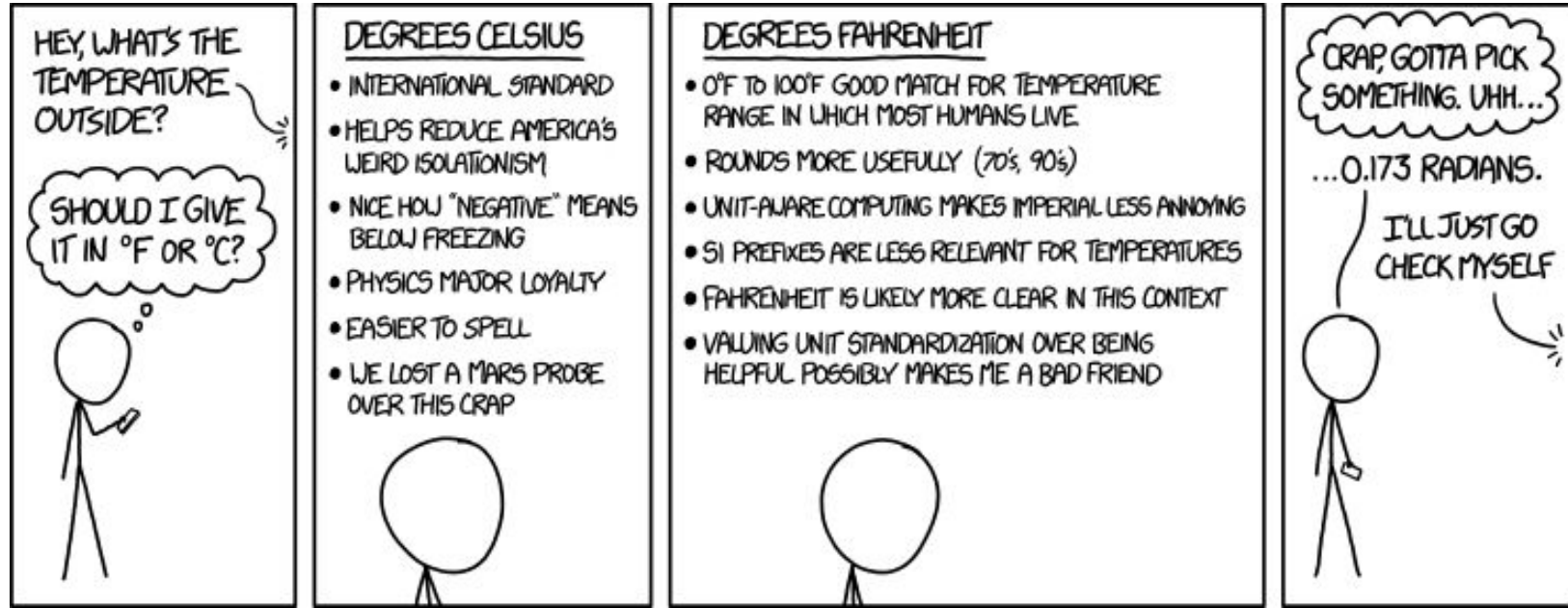
Where n is the number of spins pointing up, for instance.

# Some deference to practicalities

The kelvin scale for temperature is defined with respect to the Boltzmann constant, exactly equal to  $1.380\,649 \times 10^{-23}$

For example, through the ideal gas law by measuring the average kinetic energy, or pressure and volume.

$$\bar{p} = nkT$$



The answer Cueball gives of 0.173 radians corresponds to a geometric angle  $9.91^\circ$  ( $0.173 \times \frac{360^\circ}{2\pi}$ ). If this were "radians Celsius" it would be  $9.91^\circ \text{C}$  corresponding to  $49.8^\circ \text{F}$  and if it were "radians Fahrenheit" it would be  $9.91^\circ \text{F}$  corresponding to  $-12.3^\circ \text{C}$ . Given the temperatures in Massachusetts (where Randall lives) when this comic came out, the day after Valentine's Day 2016, Cueball was probably giving his answer in radians Fahrenheit.